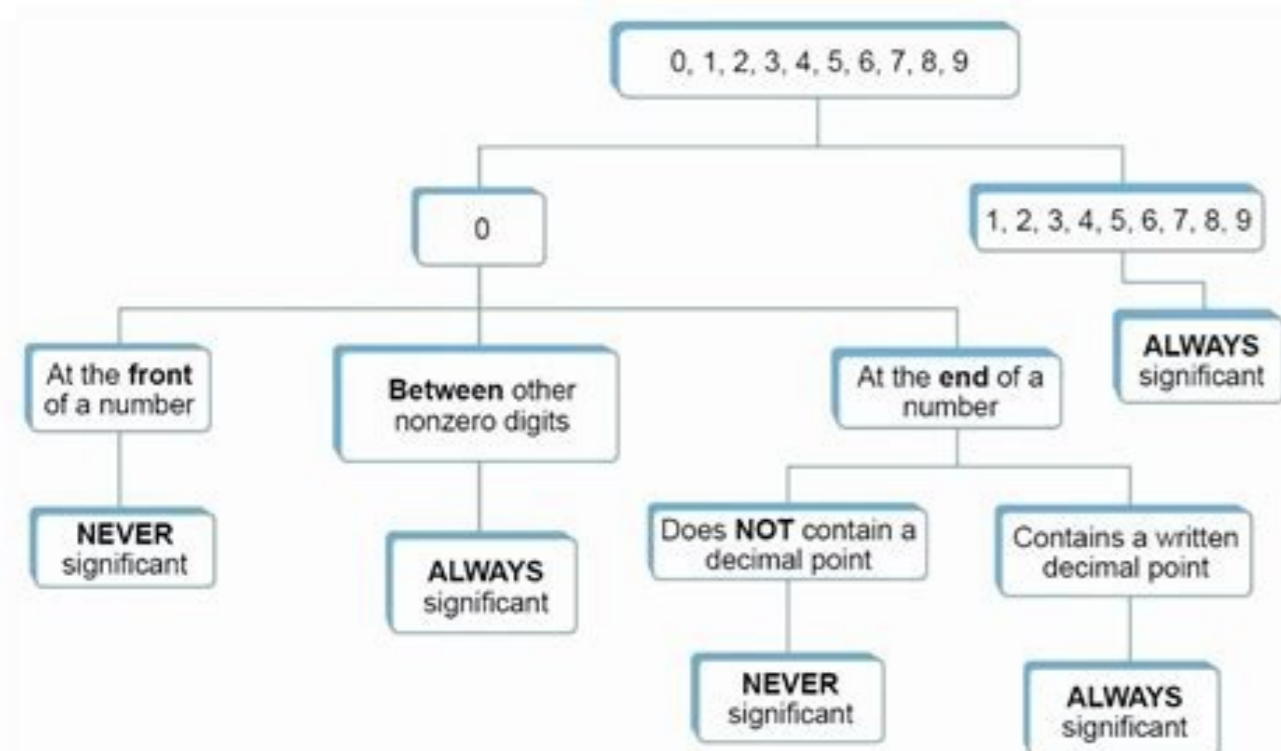
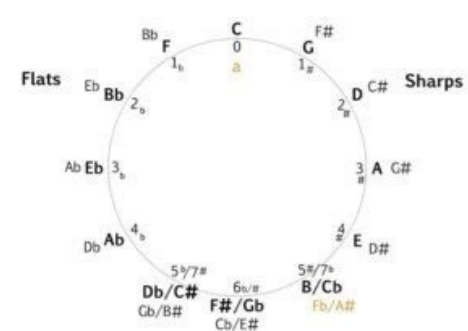


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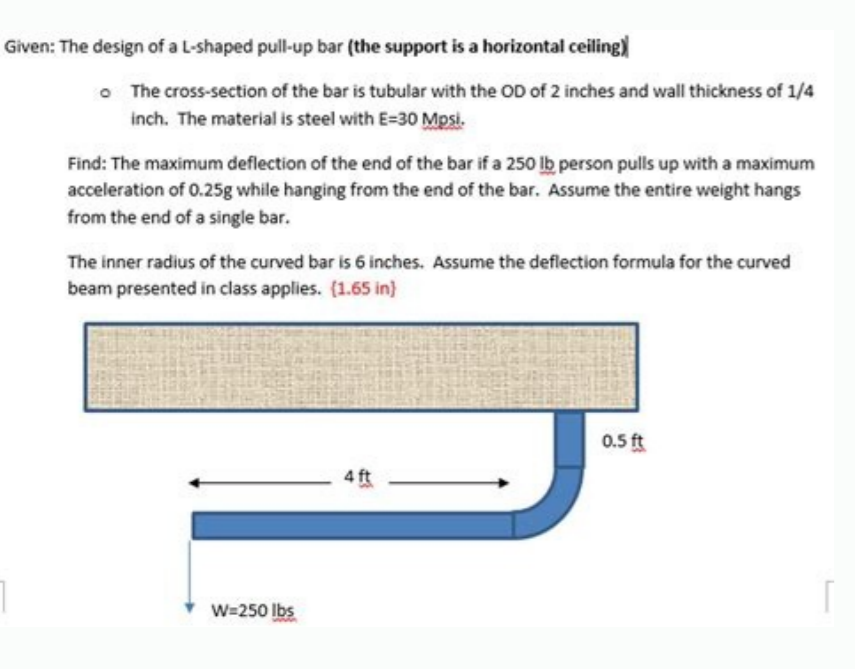


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**COMPOSITION OF HYDRATES** Name \_\_\_\_\_

A hydrate is an ionic compound with water molecules loosely bonded to its crystal lattice. The water is a specific ratio to the ionic compound. By heating, the water molecules for every one formula unit of  $\text{CuSO}_4 \cdot x\text{H}_2\text{O}$ . Answer the questions below.

1. What percentage of water is found in  $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$ ?
2. What percentage of water is found in  $\text{Fe}_2\text{SO}_4 \cdot 7\text{H}_2\text{O}$ ?
3. A  $0.5 \text{ g}$  sample of a hydrate of  $\text{NiCl}_2$  was heated, and only  $0.3 \text{ g}$  of the anhydrous salt remained. What percentage of water was in the hydrate?
4. A  $2.2 \text{ g}$  sample of a hydrate of  $\text{CuSO}_4$  was heated, and only  $1.7 \text{ g}$  of the anhydrous salt remained. What percentage of water was in the hydrate?
5. A  $3.2 \text{ g}$  sample of  $\text{Ni}_2\text{CO}_3 \cdot x\text{H}_2\text{O}$  is heated to constant mass. How much anhydrous salt remains?
6. A  $0.5 \text{ g}$  sample of  $\text{CuSO}_4 \cdot x\text{H}_2\text{O}$  is heated, and  $0.3 \text{ g}$  of the anhydrous salt remains. What is the value of  $x$ ?



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